

Reaction Stoichiometry

GUIDED NOTES

Learning Guide Reaction Stoichiometry

Name: _____
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A. Information Given by Chemical Equations

Recall that chemical changes are really arrangements of atoms that can be described by chemical equations. In this lesson, we will review the meaning and usefulness of chemical equations by considering the combustion of hydrogen gas.



The coefficients in the balanced chemical equation tell you the relative number of molecules of each substance. For example, in the reaction above, 1 molecule of O_2 is needed to react with 2 molecules of H_2 . Multiplying each coefficient by 2 gives the mole relationships between reactants and products. For example, 2 molecules of H_2 are required to react with 1 molecule of O_2 to produce 2 molecules of H_2O .

Mole Ratio

The mole ratio is a conversion factor that represents the relative number of moles of each substance. The mole ratio is derived from the coefficients of the substances involved in the reaction.

For example, the mole ratio of O_2 to H_2O is 1 mole O_2 to 2 moles H_2O .

$$\frac{2.00 \text{ mol O}_2}{1} = \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2}$$

Number of H_2O produced

$$2.00 \text{ mol O}_2 \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol O}_2} = 4.00 \text{ mol H}_2\text{O}$$

Mole-Mole Relationships

- ▶ A **mole ratio** is a conversion factor that converts the number of moles of one substance to the number of moles of another substance.
 - Used to convert between **any** two species in a chemical reaction.

Mole ratio =
coefficients in
balanced
equation



The mole ratios can be written as fractions:

$$\frac{1 \text{ mol O}_2}{2 \text{ mol H}_2\text{O}}$$

$$\frac{1 \text{ mol O}_2}{2 \text{ mol H}_2}$$

$$\frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}}$$

GUIDED NOTES ★ POWERPOINT ★ WORKSHEET

Guided Practice Activity Reaction And Revolution

B Lingard



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