

Calorimetry: Heat of Neutralization

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Manatee Community College

Purpose: Determine the calorimeter constant for a calorimeter. Determine the enthalpy of neutralization for the reaction of a strong acid and a strong base.

Background: When a system absorbs heat (q) its temperature changes by ΔT . Experimental measurements demonstrate that the heat absorbed by a system and its corresponding temperature change are directly proportional. $q=C \times \Delta T$. C is called the **heat capacity**, which is often defined as the quantity of heat required to change its temperature by 1°C .

$$C = q / \Delta T = \text{J}/^\circ\text{C}$$

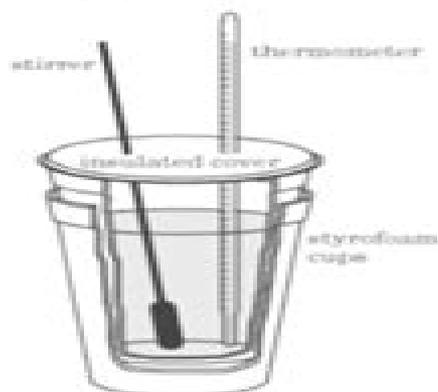
Heat capacity is an extensive property; it depends on the amount of matter being heated. The measure of intrinsic capacity of a substance to absorb heat is called **specific heat capacity** (C_s), the amount of heat required to raise the temperature of 1 gram of the substance by 1°C .

$$C_s = q / (\Delta T \cdot m)$$

The C_s of H_2O is $4.184\text{J/g}^\circ\text{C}$. For the case where all the heat is used to heat water, the amount of heat is calculated by the relation

$$q = C_s \times m_{\text{water}} \times \Delta T$$

The enthalpy change for a chemical reaction, ΔH_{rxn} is also called the **enthalpy of reaction** or **heat of reaction**, which is also an extensive property. For many aqueous reactions, ΔH_{rxn} can be measured simply using a coffee-cup calorimeter.



Heat Of Neutralization Lab Manual

David Kirk



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web für all jene die sich kompakt und leicht verständlich über bilanzen informieren wollen ist dies genau das richtige buch zunächst erklären die autoren den unterschied zwischen internem und externem rechnungswesen damit sie das thema jahresabschluss einordnen können