

Answer Key Chapter 11 Study Guide

Be able to:

describe how both ions, ionic bonds and covalent bonds form.

predict what ion will form. **Using your periodic table**

state how the bonds of a compound relate to their properties.

name compounds from the formula and write formulas from the name.

Bonding that involves the transfer of electrons is known as **Ionic Bonding**.

Bonding that involves the sharing of electrons is known as **Covalent Bonding**.

Atoms that easily lose electrons are located on the **Left side** of the periodic table. **These are the metals. Remember the nonmetals GAIN electrons in order to complete their shell because they want the magic number of 8, a stable octet.**

If an atom has a charge of +2, what does this mean? **This means that the atom has given up or lost 2 electrons so it now has 2 more protons than electrons. An example of this would be calcium. Calcium has an atomic number of 20 which means it has 20 protons and as a neutral atom has 20 electrons. When calcium becomes Ca^{+2} , it means that calcium has lost 2 electrons, in order to have a stable octet, so it now only has 18 electrons and its overall charge is more positive by 2.**

How does the group number relate to an element's valence electrons? **for groups 1 and 2, the group number IS the number of valence electrons, for groups 13-18 the valence electrons are equal to the group number minus 10.**

What types of atoms form ionic bonds? **a metal and a nonmetal (or a polyatomic ion)**

What types of atoms form covalent bonds? **2 non metals or a non metal and a metalloid**

What type of atoms normally doesn't like to bond? **the noble gases**

Glencoe Chemistry Chapter 11 Study Guide Answer

H Kauffman



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