

Dr. Sarah Ayala  
11/1/20

Date \_\_\_\_\_ Name \_\_\_\_\_  
 Section \_\_\_\_\_ Title \_\_\_\_\_  
 Instructor \_\_\_\_\_

### Pre-Lab Study Questions

- Why are burning candles and rusting nails examples of chemical change?  
 Rust is oxidation  
 Candles are combustion  
 They undergo one of the 4 criteria  
 formation of gas, formation of solid, formation  
 of an unstable, or heat produced/absorbed.
- What is included in a chemical equation?  
 a reactant and  
 a product
- How does a combination reaction differ from a decomposition reaction?  
 Combination - two elements combine to form  
 a compound.  
 Decomposition - one compound splits  
 apart.

4. Balance each of the following reactions, and identify the type of reaction.

Unbalanced Equation	Type of Reaction
a. $2Al(s) + Fe_2O_3(s) \xrightarrow{\Delta} Al_2O_3 + Fe(l)$	single replacement
b. $2KClO_3(s) \xrightarrow{\Delta} 2KCl(s) + O_2(g)$	decomposition
c. $Li(s) + Cl_2(g) \xrightarrow{\Delta} LiCl(s)$	combination
d. $C_2H_4(g) + O_2(g) \xrightarrow{\Delta} CO_2(g) + H_2O(g)$	double replacement
e. $CrCl_3(aq) + H_2S(g) \xrightarrow{\Delta} Cr_2S_3(s) + HCl(aq)$	double replacement

5. Complete and balance each of the following reactions.

Reactants	Type of Reaction
a. $K(s) + N_2(g) \longrightarrow 2KN(s)$	Combination
b. $C_3H_{12}(g) + O_2(g) \xrightarrow{\Delta}$	Combustion
c. $Al(s) + CuSO_4(aq) \longrightarrow$	Single Replacement
d. $CoCl_2(aq) + AgNO_3(aq) \longrightarrow$	Double Replacement
e. $MgCO_3(s) \xrightarrow{\Delta} MgO(s) + CO_2(g)$	Decomposition

# Lab Questions For Chapter 5

**Judd E. Hollander**



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