

CHAPTER 14: COVALENT BONDING: ORBITALS

PROBLEMS WITH LEWIS STRUCTURE

- Provides no indication of the **shapes** of the molecules
- O_2 would be predicted to be diamagnetic, but it is **paramagnetic**

Paramagnetic = has unpaired electrons.



The Lewis structure of O_2 has an $O=O$ double bond with no unpaired electrons, i.e. diamagnetic.

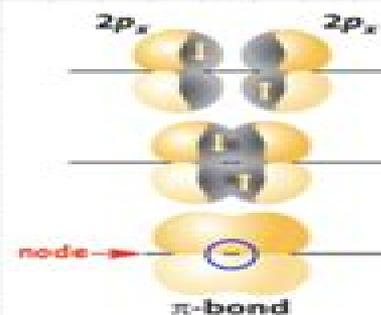
- Many compounds cannot be represented by a single structure
- Resonance structures **differ only in position of electrons** not the nuclei.
- Result in notions of hypervalent compounds and "expansion of octet"
- Can't explain molecules such as B_2H_6 (electron deficient molecules)

VALENCE-BOND THEORY "LOCALIZED ELECTRON MODEL"

- Covalent bond is created by the overlap of orbitals on two atoms

TYPES OF OVERLAP: σ and π Bonds

- **Bonding in H_2**
 - Sausage shaped orbital, bond called σ -orbital
 - Called σ orbital because it has **no node** and is **cylindrically symmetrical**
 - Electron density is **concentrated between two nuclei**, but is still highest at the nuclei themselves.
- **Bonding in N_2**
 - σ orbital is obtained by overlap of **two p_z -orbitals**
- Perpendicular p orbitals overlap in a side-by-side manner.
- The resulting orbital has two lobes of electron density and is called a **π -bond**.
- Called π -orbital because it **has a node** and, when viewed down the internuclear axis, it looks like a p-orbital
- π -bond has **no electron density between the nuclei along the bond axis**.



General Chemistry Chapter 14 Covalent Bonding Orbitals Questions

Annelies Wilder-Smith



General Chemistry Chapter 14 Covalent Bonding Orbitals Questions:

General Chemistry Chapter 14 Covalent Bonding Orbitals Questions Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**General Chemistry Chapter 14 Covalent Bonding Orbitals Questions**," compiled by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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