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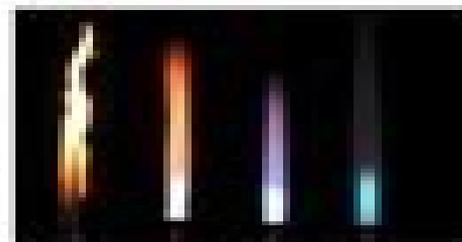
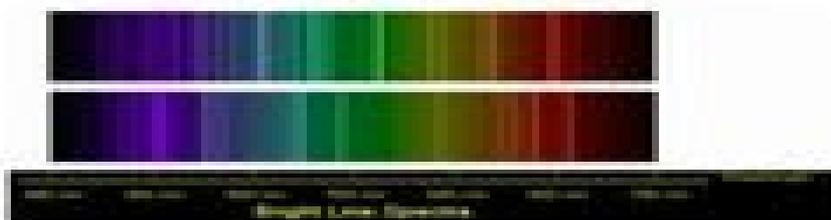
## Flame Test Lab

**Essential Questions:** What happens when different cations are exposed to a flame?

Reading Questions	Data	Claim	Evidence	Reasoning	Model
/10	/10	/10	/5	/5	/10

**Materials:**

- 1) Bunsen Burner    2) Q-tips    3) Flint lighter    4) Marquis's Reagent    5) 100 mL beaker

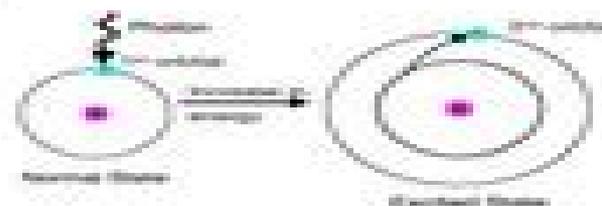


### Background Information:

If you excite an atom or an ion by very strong heating, electrons can be promoted from their normal unexcited state into higher orbitals. As they fall back down to lower levels (either in one go or in several steps), energy is released as light. Each of these jumps involves a specific amount of energy being released as light energy and each corresponds to a particular wavelength (or frequency). As a result of all these jumps, a spectrum of lines will be produced, some of which will be in the visible part of the spectrum. The color you see will be a combination of all these individual colors.

In the case of sodium (or other metal ions), the jumps involve very high energies and these result in lines in the UV part of the spectrum which your eyes can't see. The jumps that you can see in flame tests come from electrons falling from a higher to a lower level in the metal atoms.

So if, for example, you put sodium chloride which contains sodium ions, into a flame, where do the atoms come from? In the hot flame, some of the sodium ions regain their electrons to form neutral sodium atoms again. A sodium atom in an unexcited state has the structure 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>1</sup>, but within the flame there will be all sorts of excited states of the electrons. Sodium's familiar bright orange-yellow flame color results from promoted electrons falling back from the 3p<sup>1</sup> level to their normal 3s<sup>1</sup> level!



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# Flame Test And Electron Configuration Lab Answer

**Patrick Vollmar**



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