

CLASS-12 BOARDS

FORMULA SHEET

Atifa Dwaisi*

Chemistry

SOLUTIONS

Mass Percentage (w/w)% = $\frac{w_2}{w_1 + w_2} \times 100$

Volume Percentage (v/v)% = $\frac{V_2}{V_1 + V_2} \times 100$

Mass by Volume Percentage (w/v) = $\frac{w_2}{V \text{ (in ml)}} \times 100$

Mass Fraction

$X_1 = \frac{w_1}{w_1 + w_2}$ or $X_2 = \frac{w_2}{w_1 + w_2}$

Parts per million (ppm) = $\frac{w_2}{w_1 + w_2} \times 10^6$

Molarity (M) = $\frac{W_B \times 1000}{M_B \times V_{\text{soln}}}$ Unit = mol/litre

Molality (m) = $\frac{W_B \times 1000}{M_B \times W_A(g)}$ Unit = mole/kg

↳ molarity is inversely proportional to temperature

Mole Fraction (X)

$X_A = \frac{n_A}{n_A + n_B}$ $X_B = \frac{n_B}{n_A + n_B}$

$X_A + X_B = 1$ mole fraction is a unitless quantity

Henry Law

$p = K_H X_B$
↳ partial pressure of gas
↳ mole fraction of solute

RAULT'S LAW

$p_A \propto X_A$ $p_B \propto X_B$
 $p_A = p_A^0 \times X_A$ $p_B = p_B^0 \times X_B$

For Volatile Solute

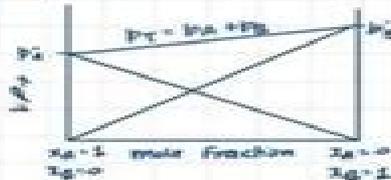
$p_r = p_A + p_B$

For Non-Volatile Solute

$p_r = p_A$
 $= p_A^0 \times X_A$

IDEAL SOLUTION

Interaction of A-B are equal to interaction of A-A and B-B



$\Delta V_{\text{mix}} = 0$
 $\Delta H_{\text{mix}} = 0$
eg n-Hexane + n-Heptane

NON-IDEAL SOLUTION

Positive Deviation

A-B interaction are weaker than A-A and B-B interaction

$\Delta V_{\text{mix}} = +ve$
 $\Delta H_{\text{mix}} = +ve$



Form minimum boiling azeotropes

eg Acetone + Ethanol

Negative Deviation

A-B interaction are stronger than A-A and B-B interaction

$\Delta V_{\text{mix}} = -ve$
 $\Delta H_{\text{mix}} = -ve$



Form maximum boiling azeotropes

eg acetone + Chloroform

Colligative Properties

depend on number of solute

Relative lowering of v.p. $\rightarrow \frac{p_A^0 - p_A}{p_A^0} = X_B$ or $i X_B$

Elevation in Boiling Point $\rightarrow \Delta T_b = K_b m$ or $i K_b m$

Depression in Freezing Point $\rightarrow \Delta T_f = K_f m$ or $i K_f m$

Osmotic Pressure $\pi = CRT$ or $\frac{n}{V} RT$ or $i \frac{W_B}{V} \times \frac{RT}{M_B}$

Unit of K_b and $K_f = K \text{ kg mol}^{-1}$

$\Delta T_b = T_b - T_b^0$

$\Delta T_f = T_f^0 - T_f$

$\Delta T_b = K_b \times \frac{W_B}{M_B} \times \frac{1000}{W_A}$

$\Delta T_f = K_f \times \frac{W_B}{M_B} \times \frac{1000}{W_A}$

ΔT_b - Elevation in B.P.T

ΔT_f - Depression in F.P.T

T_b - B.P.T of solution

T_b^0 - B.P.T of pure solvent

T_f - F.P.T of solution

K_b - molal elevation constant

K_f - molal depression constant

Van't Hoff factor (i)

= $\frac{\text{Normal Molecular Mass}}{\text{observed molecular mass}}$

$i > 1$ solute undergoes dissociation

$i < 1$ solute undergoes association

$i = 1$ no association no dissociation

% dissociation = $\frac{i-1}{n-1}$ % association = $\frac{1-i}{n-1}$

High School Chemistry Formulas Cheat Sheet

Yi-Tong Ma



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