

Honors Chemistry Unit 5 Gas Laws and States of Matter Study Guide

Directions: Use your notes (found on Schoology) to answer the following questions.

- 1) What is the Kinetic Molecular theory and what does it explain?**
States that particles in matter are always in motion. Explains each phase of matter based on its energy of particles and forces that act between them.
- 2) Does a real gas obey the Kinetic Molecular Theory?**
No.
- 3) How do the particles of an Ideal gas collide?**
They hit and bounce back (elastic).
- 4) In a gas, what determines the kinetic energy?**
Temperature and mass
- 5) What is an example of effusion?**
A nail hole in a tire.
- 6) What happens to the volume of a gas when it is compressed?**
Volume decreases as compression (pressure) increases. Seen in Boyle's Law.
- 7) Is a liquid more or less compressible than the other two common states of matter & why?**
Liquid is more compressible than a solid but way less compressible than a gas.
- 8) Intermolecular forces are the strongest in what state of matter?**
Solids
- 9) List solids, liquids, and gases in order of increasing energy.**
solids < liquids < gases
- 10) Why are solids and liquids denser than gases?**
They are at lower temperatures so the particles move slower and are closer together.
- 11) What is the difference between amorphous and crystalline solids?**
Crystalline solids have orderly repeating geometric patterns. Amorphous solids do not (example plastics and glass).
- 12) Define triple point.**
All three phases of matter coexist
- 13) Define critical temperature.**
The point where a substance can only exist as a gas past that temperature.
- 14) What is it called when a substance changes directly from a solid to a gas?**
Sublimation
- 15) What is it called when a substance skips the liquid phase and goes straight from a gas to a solid?**
Deposition
- 16) When the pressure is greater than 1 atm, how would that affect the temperature at which water boils?**
It would increase the surface tension causing the temperature need to boil the water to increase.
- 17) When undergoing a phase change, the temperature is constant.**
- 18) Convert the following: $125^{\circ}\text{C} = 398\text{ K}$, $286\text{K} = 13^{\circ}\text{C}$, $124\text{K} = -149^{\circ}\text{C}$**

Gas Laws Honors Study Guide

N Colangelo



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