

## Passage I

Below is a table showing some Standard State Reduction Potentials.

Half reaction	E (V)
$F_2 + 2e^- \rightarrow 2F^-$	2.87
$Cl_2 + 2e^- \rightarrow 2Cl^-$	1.36
$Br_2 + 2e^- \rightarrow 2Br^-$	1.09
$Ag^+ + e^- \rightarrow Ag$	0.80
$Fe^{3+} + e^- \rightarrow Fe^{2+}$	0.77
$Cu^{2+} + 2e^- \rightarrow Cu$	0.34
$2H^+ + 2e^- \rightarrow H_2$	0.00
$Pb^{2+} + 2e^- \rightarrow Pb$	-0.13
$Ni^{2+} + 2e^- \rightarrow Ni$	-0.25
$Zn^{2+} + 2e^- \rightarrow Zn$	-0.76
$Na^+ + e^- \rightarrow Na$	-2.71

- What should happen when a piece of copper is placed in 1M HCl?
  - The copper is dissolved by the acid.
  - The copper is dissolved by the acid with the release of hydrogen gas.
  - The copper bursts into greenish flames.
  - Nothing happens.
- What should happen when a piece of lead is placed in 1M HCl?
  - The lead is dissolved by the acid.
  - The lead begins to dissolve with the release of hydrogen gas.
  - The lead bursts into flames.
  - Nothing happens.
- How many voltaic cells with a voltage greater than 2V can be made using the standard half-cell reactions listed?
  - 11
  - 9
  - 17
  - 15
- Suppose that the half-cell reaction  $Ni^{2+} + 2e^- \rightarrow Ni$  is assigned to have a potential of 0.00 V. The standard reduction potential for the reduction of  $Br_2$  would be?
  - 1.34 V
  - 0.84 V
  - 1.09 V
  - 1.25 V

- A voltaic cell is set up with  $F_2/F^-$  as one half cell and  $Br_2/Br^-$  as the other. What is the voltage of the cell at standard state?
  - 3.96 V
  - 1.78 V
  - 1.09 V
  - 2.87 V
- $NaCl$  is melted in a beaker and electrolysis carried out. What is the minimum potential difference required to separate sodium and chlorine?
  - 4.07 V
  - 1.36 V
  - 1.36 V
  - 4.07 V
- If standard state oxidation potentials are used instead of standard state reduction potentials and the half reactions are listed in descending order, which one of the following would be true?
  - the lead reaction would be above the silver reaction
  - the fluorine reaction would be above the chlorine reaction
  - the iron reaction would be above the nickel reaction
  - the hydrogen reaction would be below the copper reaction

## Passage II

A student is provided with a variety of electrical components and asked to perform certain tasks.

- She is provided with three 2  $\Omega$  resistors. How many different total resistances can be produced if all the resistors must be used in a circuit?
  - 3
  - 4
  - 5
  - 6
- What are the minimum and maximum total resistances produced above?
  - 0.67  $\Omega$  and 3  $\Omega$
  - 1  $\Omega$  and 6  $\Omega$
  - 1.33  $\Omega$  and 6  $\Omega$
  - 0.67  $\Omega$  and 6  $\Omega$
- What is the total resistance of two 2  $\Omega$  resistors in parallel connected to a third 2  $\Omega$  resistor in series?
  - 6
  - 1
  - 3
  - 1.33

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**Xiang Xie**



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