

## Observations

	1	Distilled H <sub>2</sub> O	--	non
HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub> (aq): acetic acid	2	Tap water	--	weak
	3	NaCl	5	strong
	4	CuSO <sub>4</sub>	5	strong
HCl(aq): hydrochloric acid	5	Sugar	5	non
	6	Alcohol	5	non
NaOH(aq): sodium hydroxide	7	HCl	5	strong
	8	HCl	1	strong
CuSO <sub>4</sub> (aq): copper(II) sulfate	9	HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	5	weak
	10	HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	1	weak
	11	NaOH	5	strong
	12	NaOH	1	strong
	13	HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	17.4	non

## Answers

1. An electrolyte is a compound that, in aqueous solution, conducts electricity. As liquids: salts, bases. When aq: salts, acids, bases.
2. HCl(aq): H<sup>+</sup>(aq) + Cl<sup>-</sup>(aq)  
NaOH(aq): Na<sup>+</sup>(aq) + OH<sup>-</sup>(aq)  
CuSO<sub>4</sub>(aq): Cu<sup>2+</sup>(aq) + SO<sub>4</sub><sup>2-</sup>(aq)
3. Acid: a compound that ionizes in water to form hydrogen ions. Base: an ionic hydroxide that dissociates in water to produce hydroxide ions

# Lab Conductivity Quizz With Answers

**Ying-Ying Zheng**



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