

Intro to Energy, Enzymes, and Metabolism

Sept 24th

Concepts in Energy and Thermodynamics

- Energy
- Reactions
- Entropy
- 2nd law of thermodynamics
- Free energy
- Kinetics of chemical reactions
- ATP
- Catalysis and Enzymes

Energy is the ability to do work, and it will determine the **direction** and **rate** of reaction ($aA + bB \leftrightarrow cC + dD$)

Enthalpy (H)

- The total energy of a system

Free Energy (G)

- The amount of a systems energy that is available and can be used to promote change or do

Entropy (S)

- The degree of disorder of a system
- Higher Entropy → More Disordered
- Lower Entropy → Highly Ordered

T → Temperature

Total Energy = Usable Energy + Unusable Energy

$$(H = G + TS)$$

$$(\Delta G = \Delta H - T \Delta S)$$

- Negative Delta G → loss of energy = Spontaneous reaction

2nd Law of Thermodynamics

- The state of a natural system itself can be reserved, but **not** without increasing the entropy of the systems surroundings
- The change in free energy determines reaction direction
- Spontaneous doesn't need a input of energy
- Delta G will tell you what direction it will go, not the rate of the reaction

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Table of Contents Mastering Biology Ch6

1. Understanding the eBook Mastering Biology Ch6
 - The Rise of Digital Reading Mastering Biology Ch6
 - Advantages of eBooks Over Traditional Books
2. Identifying Mastering Biology Ch6
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Biology Ch6
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mastering Biology Ch6
 - Personalized Recommendations
 - Mastering Biology Ch6 User Reviews and Ratings
 - Mastering Biology Ch6 and Bestseller Lists
5. Accessing Mastering Biology Ch6 Free and Paid eBooks
 - Mastering Biology Ch6 Public Domain eBooks
 - Mastering Biology Ch6 eBook Subscription Services
 - Mastering Biology Ch6 Budget-Friendly Options
6. Navigating Mastering Biology Ch6 eBook Formats
 - ePub, PDF, MOBI, and More
 - Mastering Biology Ch6 Compatibility with Devices
 - Mastering Biology Ch6 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mastering Biology Ch6
 - Highlighting and Note-Taking Mastering Biology Ch6
 - Interactive Elements Mastering Biology Ch6
8. Staying Engaged with Mastering Biology Ch6

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Biology Ch6
9. Balancing eBooks and Physical Books Mastering Biology Ch6
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Biology Ch6
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mastering Biology Ch6
- Setting Reading Goals Mastering Biology Ch6
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mastering Biology Ch6
- Fact-Checking eBook Content of Mastering Biology Ch6
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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