

Name: **Brittany French**

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Student Exploration: Electron Configuration

Activity B: Atomic radii	<u>Get the Gizmo ready:</u> <ul style="list-style-type: none"> • On the PERIODIC TABLE tab, select Na (sodium). • Select the ELECTRON CONFIGURATION tab. 	
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Question: How do the radii of atoms change across a period of the periodic table?

1. **Predict:** Positively charged protons in the nucleus of an atom are attracted to negatively charged electrons.

How do you think the **atomic radii** will change as electrons are added to a shell?

It will decrease

2. **Arrange:** Create a proper electron configuration for sodium. After clicking **Check**, note the **Electron configuration** and the **Atomic radius** now listed at right.

Sodium electron configuration: **1s² 2s² 2p⁶ 3s¹** Atomic radius: **190 Pico meters**

3. **Compare:** Click **Next element**, and then add an electron to the magnesium atom. Click **check**, and record the electron configuration and atomic radius below.

Magnesium electron configuration: **1s² 2s² 2p⁶ 3s²** Atomic radius: **145 Pico meters**

4. **Gather data:** Create electron configurations for the next six elements. Record the electron configuration and atomic radius of each. (Note: The symbol for picometer is *pm*.)

Element	Number of electrons	Electron configuration	Atomic radius (pm)
Aluminum	13	1s ² 2s ² 2p ⁶ 3s ² 3p ¹	118
Silicon	14	1s ² 2s ² 2p ⁶ 3s ² 3p ²	111
Phosphorus	15	1s ² 2s ² 2p ⁶ 3s ² 3p ³	98
Sulfur	16	1s ² 2s ² 2p ⁶ 3s ² 3p ⁴	88
Chlorine	17	1s ² 2s ² 2p ⁶ 3s ² 3p ⁵	79
Argon	18	1s ² 2s ² 2p ⁶ 3s ² 3p ⁶	71

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