

Activity A:
Small atoms

Get the Gizmo ready:

- On the PERIODIC TABLE tab, select H (hydrogen).
- Select the ELECTRON CONFIGURATION tab.
- Click **Reset**.



Introduction: Electrons are arranged in **orbitals**, **subshells**, and **shells**. These levels of organization are shown by the boxes of the Gizmo. Each box represents an orbital. The subshells are labeled with letters (*s*, *p*, *d*, and *f*) and the shells are labeled with numbers.

Question: How are electrons arranged in elements with atomic numbers 1 through 10?

1. **Infer:** Based on its atomic number, how many electrons does a hydrogen atom have? _____
2. **Arrange:** The **Aufbau principle** states that electrons occupy the lowest-energy orbital. Click once in the **1s** box to add an electron to the only orbital in the *s* subshell of the first shell.
Click **Check**. What is the electron configuration of hydrogen? _____
3. **Arrange:** Click **Next element** to select helium. Add another electron to the **1s** orbital. The arrows represent the **spin** of the electron. What do you notice about the arrows?

The **Pauli exclusion principle** states that electrons sharing an orbital have opposite spins.

4. **Check your work:** Click **Check**. What is the electron configuration of helium? _____
5. **Arrange:** Click **Next element** and create electron configurations for lithium, beryllium, and boron. Click **Check** to check your work, and then list each configuration below:

Lithium: _____ Beryllium: _____ Boron: _____

6. **Arrange:** Click **Next element** to select carbon. Add a second electron to the first **2p** orbital.
Click **Check**. What feedback is given? _____

7. **Rearrange:** **Hund's rule** states that electrons will occupy an empty orbital when it is available in that subshell. Rearrange the electrons within the **2p** subshell and click **Check**.

Is the configuration correct now? _____

1s

Show the correct configuration in the boxes at right:

2s 2p

(Activity A continued on next page)

Explore Learning Element Builder Answer Key

LL Leslie

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