

13. Consider the following statements
- Atmospheric pressure is less at the top of a mountain than at sea level
  - Gases are much more compressible than solids or liquids
  - When the atmospheric pressure increases the height of the mercury column rises
- Select the correct statement
- I and II
  - II and III
  - I and III
  - I, II and III
14. Compressibility factor for  $\text{CO}_2$  at 400 K and 71.0 bar is 0.8697. The molar volume of  $\text{CO}_2$  under these conditions is
- 22.04  $\text{dm}^3$
  - 2.24  $\text{dm}^3$
  - 0.41  $\text{dm}^3$
  - 19.5  $\text{dm}^3$
15. If temperature and volume of an ideal gas is increased to twice its values, the initial pressure  $P$  becomes
- 4P
  - 2P
  - P
  - 3P
16. At identical temperature and pressure, the rate of diffusion of hydrogen gas is  $3\sqrt{3}$  times that of a hydrocarbon having molecular formula  $\text{C}_n\text{H}_{2n-2}$ . What is the value of  $n$ ?
- 8
  - 4
  - 3
  - 1
17. Equal moles of hydrogen and oxygen gases are placed in a container, with a pin-hole through which both can escape what fraction of oxygen escapes in the time required for one-half of the hydrogen to escape. (NEET phase I)
- $\frac{3}{8}$
  - $\frac{1}{2}$
  - $\frac{1}{8}$
  - $\frac{1}{4}$
18. The variation of volume  $V$ , with temperature  $T$ , keeping pressure constant is called the coefficient of thermal expansion i.e.  $\alpha = \frac{1}{V} \left( \frac{\partial V}{\partial T} \right)_P$ . For an ideal gas  $\alpha$  is equal to
- T
  - $\frac{1}{T}$
  - P
  - none of these
19. Four gases P, Q, R and S have almost same values of 'b' but their 'a' values (a, b are Vander Waals Constants) are in the order  $Q < R < S < P$ . At a particular temperature, among the four gases the most easily liquefiable one is
- P
  - Q
  - R
  - S
20. Maximum deviation from ideal gas is expected from (NEET)
- $\text{CH}_4(\text{g})$
  - $\text{NH}_3(\text{g})$
  - $\text{H}_2(\text{g})$
  - $\text{N}_2(\text{g})$
21. The units of Vander Waals constants 'b' and 'a' respectively
- $\text{mol L}^{-1}$  and  $\text{L atm}^2 \text{mol}^{-2}$
  - $\text{mol L}$  and  $\text{L atm mol}^2$
  - $\text{mol}^{-1} \text{L}$  and  $\text{L}^2 \text{atm mol}^{-2}$
  - none of these

# Gaseous State Multiple Choice Questions

**Joacim Rocklöv**



**Gaseous State Multiple Choice Questions:**

## Enjoying the Melody of Expression: An Psychological Symphony within **Gaseous State Multiple Choice Questions**

In a world taken by monitors and the ceaseless chatter of instantaneous connection, the melodic splendor and psychological symphony produced by the prepared word usually fade in to the backdrop, eclipsed by the relentless noise and distractions that permeate our lives. But, situated within the pages of **Gaseous State Multiple Choice Questions** a wonderful fictional prize overflowing with fresh feelings, lies an immersive symphony waiting to be embraced. Crafted by a masterful composer of language, this fascinating masterpiece conducts readers on an emotional journey, skillfully unraveling the hidden tunes and profound affect resonating within each cautiously crafted phrase. Within the depths with this moving examination, we shall discover the book is key harmonies, analyze their enthralling writing type, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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