

## Practice with the Gas Laws with Answers

All temperatures must be in **KELVIN** STP = **0 °C and 1.0 atm**

Converting one unit of pressure to another

1. Change 758. mm Hg to Torr

$$758 \text{ mm Hg} \times \frac{760 \text{ Torr}}{760 \text{ mm Hg}} = 758 \text{ Torr}$$

2. Change 771.3 mm Hg to atm

$$771.3 \text{ mm Hg} \times \frac{1 \text{ atm}}{760 \text{ mm Hg}} = 1.015 \text{ atm}$$

3. Change 95.6 kPa to mm Hg

$$95.6 \text{ kPa} \times \frac{760 \text{ mm Hg}}{101.325 \text{ kPa}} = 717 \text{ mm Hg}$$

4. Is the pressure of the gas sample greater or less than the pressure of the atmosphere? Explain.

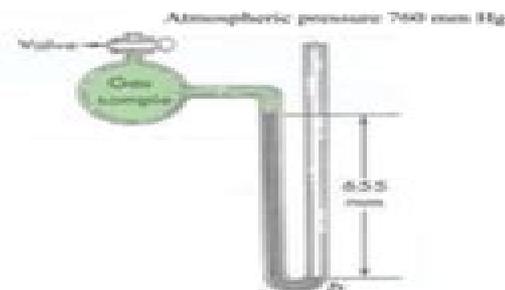
**Less, the mercury line is higher on that side**

What is the pressure of the gas in this manometer (in mm Hg)?

$$\begin{array}{r} 760 \text{ mm Hg} \\ -655 \text{ mm Hg} \\ \hline 105 \text{ mm Hg} \end{array}$$

Convert that pressure to kPa.

$$105 \text{ mm Hg} \times \frac{101.325 \text{ kPa}}{760 \text{ mm Hg}} = 14.0 \text{ kPa}$$



Calculating Density of a Gas at STP

1. What is the density of dinitrogen tetroxide at STP?

$$\text{N}_2\text{O}_4 \quad \frac{92.01 \text{ g}}{22.4 \text{ L}} = 4.108 \text{ g/L}$$

2. What is the density of sulfur trioxide at STP?

$$\text{SO}_3 \quad \frac{80.06 \text{ g}}{22.4 \text{ L}} = 3.574 \text{ g/L}$$

# Gas Laws Practice Problems With Solutions

**RJ Alexander**



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