

# Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table

## Chemistry

### LIMITING REACTANTS

1. The reaction below produces ammonia for use in fertilizers.  
 $N_2 + 3H_2 \rightarrow 2NH_3$   
If you start with 28 grams of  $N_2$  and 12 grams of  $H_2$ , determine the limiting reactant.

2. Calculate the grams of ammonia produced.

3. How many grams of the excess reactant is present in a left over?  
\* Multiply the mass of the limiting reactant by the mole ratio of the excess reactant.

4. Calculate moles of product \* mole of reactant.

2. The reaction below produces carbon dioxide and water from a common household product. Also, carbon dioxide is used in many applications.  
 $2C_2H_5OH + O_2 \rightarrow 2CO_2 + 2H_2O$   
If you have 100 grams of ethyl alcohol and 50 grams of oxygen, determine the limiting reactant.

3. How many grams of water produced are produced?



4. Magnesium chloride is used to create salts for de-icing roads. This reacts with any magnesium chloride is produced.  
 $Mg + 2HCl \rightarrow MgCl_2 + H_2$   
If you start with 20 grams of magnesium and 20 grams of hydrochloric acid, determine the limiting reactant.

5. Calculate the grams of magnesium chloride produced.

6. How many grams of the excess reactant is present in a left over?

7. The reaction below is used to create water for drinking. This reacts with any hydrogen chloride is produced.  
 $Ca + 2HCl \rightarrow CaCl_2 + H_2$   
If you have 100 grams of calcium and 50 grams of hydrochloric acid, determine the limiting reactant.

8. How much carbon dioxide is produced?

9. The reaction below produces carbon dioxide and water from a common household product. Also, carbon dioxide is used in many applications.  
 $2C_2H_5OH + O_2 \rightarrow 2CO_2 + 2H_2O$   
If you have 100 grams of ethyl alcohol and 50 grams of oxygen, determine the limiting reactant.

10. Calculate the grams of water produced.

11. How many grams of the excess reactant is present in a left over?  
\* Multiply the mass of the limiting reactant by the mole ratio of the excess reactant.

12. Calculate moles of product \* mole of reactant.

# ANSWER KEY INCLUDED

# Limiting Reactant Problems Answer Key

**RC Schank**



### **Limiting Reactant Problems Answer Key:**

*Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

*Mapping College Chemistry* Stephen DeMeo, 2019-03-01 This text is a chemistry problem solving resource appropriate for teachers and their students who are enrolled in high school Advanced Placement Chemistry or in a first year college General Chemistry course The book incorporates a chemistry problem solving plan one that uses an innovative graphic organizer strategy The strategy successfully evaluated with students combines problem solving processes with chemical concepts that will allow students to solve the most common and difficult problems encountered in the first year of chemistry Topical problem solving will focus on limiting reactant stoichiometry identifying types of chemical reactions equilibrium acid base equilibria and electrochemistry Why would this resource be of interest to chemistry students To be successful to get into a well known college medical school physical therapy or graduate program often requires that students get an A in your pre requisite Introductory General Chemistry course To make matters worse many college professors feel that only a few students should get A grades and therefore they give difficult exams that many students fail this is the weeding out process that every pre health student is apprehensive about To succeed in this competitive environment entails not just studying harder or longer it means re organizing textbook content so that it is meaningful to the student This is the first text of its kind to employ a reliable research based strategy that incorporates a decision based visual tool to solve chemistry textbook problems ones that can make or break a career

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