

Name Mel

Locker # _____

Bonding & Molecular Structure**LEWIS STRUCTURES**

Indicate the total # of VALENCE electrons for each molecule (VE's). Calculate the number of bonded electrons from the skeleton (BE's). Calculate the Remaining Electrons (RE's = VE's - BE's) and distribute the remaining electrons onto the skeletal structure. Finally, satisfy the octets by creating double and triple bonds. Your final structures should be given below and your scratch paper should clearly show your work.

F_2 VE's = <u>14</u> BE's = <u>2</u> RE's = <u>12</u> $:\ddot{F}-\ddot{F}:$ Shape: <u>linear</u>	H_2 VE's = <u>2</u> BE's = <u>2</u> RE's = <u>0</u> $H-H$ Shape: <u>linear</u>	HF VE's = <u>8</u> BE's = <u>2</u> RE's = <u>6</u> $H-\ddot{F}:$ Shape: <u>linear</u>	NH_3 VE's = <u>8</u> BE's = <u>6</u> RE's = <u>2</u> $H-\overset{\cdot\cdot}{\underset{\cdot\cdot}{N}}-H$ Shape: <u>trigonal pyramidal</u>
CH_4 VE's = <u>8</u> BE's = <u>8</u> RE's = <u>0</u> $H-\overset{\cdot\cdot}{\underset{\cdot\cdot}{C}}-H$ Shape: <u>tetrahedral</u>	C_2H_6 VE's = <u>14</u> BE's = <u>14</u> RE's = <u>0</u> $H-\overset{\cdot\cdot}{\underset{\cdot\cdot}{C}}-\overset{\cdot\cdot}{\underset{\cdot\cdot}{C}}-H$ Shape: <u>tetrahedral (both)</u>	AlH_3 VE's = <u>6</u> BE's = <u>6</u> RE's = <u>0</u> $H-Al-H$ Shape: <u>trigonal planar</u>	BH_3 VE's = <u>6</u> BE's = <u>6</u> RE's = <u>0</u> $H-\overset{\cdot\cdot}{\underset{\cdot\cdot}{B}}-H$ Shape: <u>trigonal planar</u>
C_2H_4 VE's = <u>12</u> BE's = <u>12</u> RE's = <u>0</u> $H-\overset{\cdot\cdot}{\underset{\cdot\cdot}{C}}=\overset{\cdot\cdot}{\underset{\cdot\cdot}{C}}-H$ Shape around one C: <u>trigonal planar</u>	C_2F_4 VE's = <u>36</u> BE's = <u>12</u> RE's = <u>24</u> $:\ddot{F}-\overset{\cdot\cdot}{\underset{\cdot\cdot}{C}}=\overset{\cdot\cdot}{\underset{\cdot\cdot}{C}}-\ddot{F}:$ Shape around one C: <u>trigonal planar</u>	CO VE's = <u>10</u> BE's = <u>6</u> RE's = <u>4</u> $\overset{\cdot\cdot}{C}\equiv\overset{\cdot\cdot}{O}$ Shape: <u>linear</u>	O_2 VE's = <u>12</u> BE's = <u>4</u> RE's = <u>8</u> $:\overset{\cdot\cdot}{O}=\overset{\cdot\cdot}{O}:$ Shape: <u>linear</u>

Modeling Lewis Dot Structures Answer Key

Steven S. Zumdahl



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An Introduction to Chemistry Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the topic of introductory chemistry in detail with specific references to examples of topics in common or everyday life It provides a major overview of topics typically found in first year chemistry courses in the USA The textbook is written in a conversational question based format with a well defined problem solving strategy and presented in a way to encourage readers to think like a chemist and to think outside of the box Numerous examples are presented in every chapter to aid students and provide helpful self learning tools The topics are arranged throughout the textbook in a traditional approach to the subject with the primary audience being undergraduate students and advanced high school students of chemistry

[The Educational Testing Act of 1981](#) United States. Congress. House. Committee on Education and Labor. Subcommittee on Elementary, Secondary, and Vocational Education, 1982

Chemistry Steven S. Zumdahl, 1989

Separation Process Principles J. D. Seader, Ernest J. Henley, D. Keith Roper, 2023-03-21 Separation Process Principles with Applications Using Process Simulator 4th Edition is the most comprehensive and up to date treatment of the major separation operations in the chemical industry The 4th edition focuses on using process simulators to design separation processes and prepares readers for professional

practice *Chemical Principles* Steven S. Zumdahl,1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

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