

## Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H<sub>2</sub>O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	<b>58.5 g</b>

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	<b>84.0</b>

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	<b>159.6</b>

# Modeling Instruction Answer Key

**London univ, exam. papers**



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## **Table of Contents Modeling Instruction Answer Key**

1. Understanding the eBook Modeling Instruction Answer Key
  - The Rise of Digital Reading Modeling Instruction Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Instruction Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Instruction Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Instruction Answer Key

- Personalized Recommendations
  - Modeling Instruction Answer Key User Reviews and Ratings
  - Modeling Instruction Answer Key and Bestseller Lists
5. Accessing Modeling Instruction Answer Key Free and Paid eBooks
    - Modeling Instruction Answer Key Public Domain eBooks
    - Modeling Instruction Answer Key eBook Subscription Services
    - Modeling Instruction Answer Key Budget-Friendly Options
  6. Navigating Modeling Instruction Answer Key eBook Formats
    - ePub, PDF, MOBI, and More
    - Modeling Instruction Answer Key Compatibility with Devices
    - Modeling Instruction Answer Key Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Modeling Instruction Answer Key
    - Highlighting and Note-Taking Modeling Instruction Answer Key
    - Interactive Elements Modeling Instruction Answer Key
  8. Staying Engaged with Modeling Instruction Answer Key
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Modeling Instruction Answer Key
  9. Balancing eBooks and Physical Books Modeling Instruction Answer Key
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Modeling Instruction Answer Key
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Modeling Instruction Answer Key
    - Setting Reading Goals Modeling Instruction Answer Key
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Modeling Instruction Answer Key

- Fact-Checking eBook Content of Modeling Instruction Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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