

Limiting Reactant Practice

5 pages + 5-page Answer KEY + Periodic Table



LIMITING REACTANTS

1. The reaction below produces ammonia for use in fertilizers.
 $N_2 + 3H_2 \rightarrow 2NH_3$
If you start with 28 grams of N_2 and 12 grams of H_2 , determine the limiting reactant.

2. Calculate the grams of ammonia produced.

3. How many grams of the excess reactant is produced in a left over?
* Multiply the mass of the limiting reactant by the mole ratio of the excess reactant.

4. Calculate moles of product * mole of reactant.

2. The reaction below produces carbon dioxide and water for a gas.
 $C_2H_4 + 3O_2 \rightarrow 2CO_2 + 2H_2O$
If you have 10 grams of C_2H_4 and 40 grams of O_2 , determine the limiting reactant.

3. How many grams of water produced are produced?



4. Magnesium chloride is used to create water for drinking tanks. The tank will only magnesium chloride is produced.
 $Mg + 2HCl \rightarrow MgCl_2 + H_2$
If you start with 20 grams of magnesium and 20 grams of hydrochloric acid, determine the limiting reactant.

5. Calculate the grams of magnesium chloride produced.

6. How many grams of the excess reactant is produced in a left over?

7. The reaction below is used to create water for drinking tanks.
 $2H_2 + O_2 \rightarrow 2H_2O$
If you have 20 grams of hydrogen and 20 grams of oxygen, determine the limiting reactant.

8. How many grams of water is produced?

9. The reaction below is used to create water for drinking tanks.
 $2H_2 + O_2 \rightarrow 2H_2O$
If you have 20 grams of hydrogen and 20 grams of oxygen, determine the limiting reactant.

10. Calculate the grams of water produced.

11. How many grams of the excess reactant is produced in a left over?

12. Calculate the grams of water produced.

13. How many grams of the excess reactant is produced in a left over?

14. Calculate the grams of water produced.

15. How many grams of the excess reactant is produced in a left over?

ANSWER KEY INCLUDED

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Howard Stephen Stoker



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