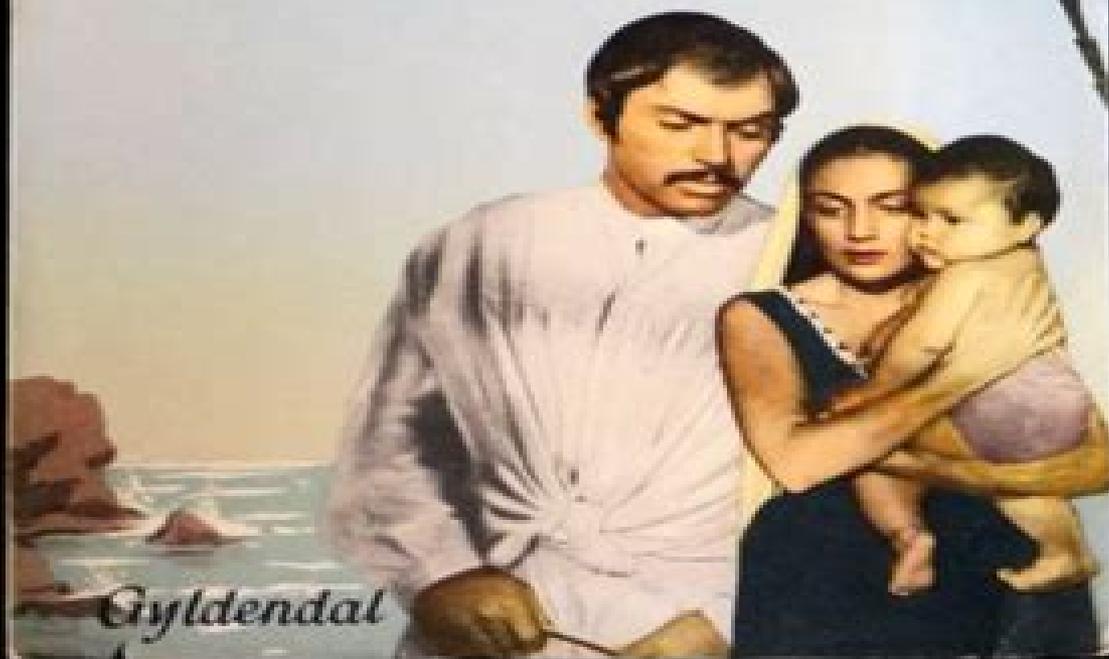


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# Perlen



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**Marcel A. Müller**



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web log H<sup>+</sup> = 10<sup>-4.05</sup> = 8.91 × 10<sup>-5</sup> mol dm<sup>-3</sup> log pK<sub>a</sub> = pH - 3.75 = 4.05 - 0.30 = 10.0 - 3.0 = 7.0 and HCOOH = 0.1 so HCOO<sup>-</sup> = 0.200 mol dm<sup>-3</sup> amount vol x

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web only the better candidates could write successive ionisation equations in part c while part d about the pH of equimolar strong and weak acid solutions was poorly answered

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web section 2 pH of strong acids number of protons released monoprotic acid acid that releases one H<sup>+</sup> ion per molecule e.g.

hcl hydrochloric acid hno<sub>3</sub> nitric acid ch

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