

Mayya's Equilibrium Study Guide



K_c is the Equilibrium Constant

$$K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$



K_p is the Equilibrium Constant
when pressures are given

$$K_p = \frac{P_c^c P_d^d}{P_a^a P_b^b}$$

$$K_p = K_c (RT)^{\Delta n}$$

Where $R=0.08206 \text{ L}\cdot\text{atm}/\text{mol}\cdot\text{K}$, T = temperature in K,
 Δn = moles of gas products - moles of gas reactants

How To Predict The Direction Of Reaction

$$Q = \frac{[C]^c \text{ noneq} [D]^d \text{ noneq}}{[A]^a \text{ noneq} [B]^b \text{ noneq}}$$

If $Q > K$, the reaction will go to the left
If $Q < K$, the reaction will go to the right
If $Q = K$, the reaction is at equilibrium

ICE CHART

I = Initial (initial concentrations given in the problem)
C = Change (in terms of X, including coefficients)
E = Equilibrium (I+C)

Make an ICE Chart when you are asked to find equilibrium concentrations or pressures

Equation Stoichiometry & Equilibrium Constant

- Reverse an equation = Take reciprocal of K ($1/K$)
- Multiply an equation by a constant = Raise K to the power of that constant
- Add equations = Multiply their K 's

Guide Chemical Equilibrium

S Ben Porath



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