

Physical changes and chemical reactions

Physical changes	Chemical reactions
Do not make new substances.	Always make one or more new substances.
Are often easy to reverse.	Are usually difficult to reverse.
The substances may change state or just be mixed together.	The new substances have different properties from the original substances.
Examples include: melting, boiling, condensing, freezing.	Examples include: combustion, neutralisation, thermal decomposition.

Gas pressure

Gas pressure is caused by the force of the particles hitting the walls of the container.

Change that increases pressure	Reason
increase the temperature	the particles move faster and so hit the walls of the container with more force and more often
increase the number of particles in the container	the particles are closer together and hit the walls of the container more often
decrease the volume of the container	the particles are closer together and hit the walls of the container more often

The reactivity series

This is a list of metals in order of reactivity, with the most reactive at the top.

The metals that react with water produce a metal hydroxide and hydrogen.

The metals that react with dilute acids produce a salt and hydrogen.

Most metals react with oxygen from the air to form metal oxides. This is an **oxidation** reaction.

Rusting of iron

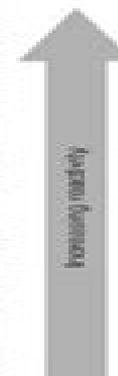
Steel is an alloy containing iron mixed with small amounts of carbon and sometimes other metals. Iron and steel need air and water to rust. Salt makes them rust more quickly than usual.

Rusting can be prevented by:

- a physical barrier to stop the air and water being in contact with the iron
- sacrificial protection, in which blocks of a more reactive metal, such as zinc or magnesium, are attached to the iron. They then corrode instead of the iron.

Stainless steel is an alloy of iron containing chromium and it does not rust.

Metal	Reaction with oxygen in air	Reaction with cold water	Reaction with dilute acid
potassium			
sodium		///	
lithium		///	///
calcium		///	///
magnesium		/	///
aluminium	///	•••	///
zinc	///	•••	///
iron	///	•••	/
tin	/	•••	/
lead	/	•••	/
copper	/	X	X
mercury	•••	X	X
silver	•••	X	X
gold	X	X	X
platinum	X	X	X



Key

	explosive		can catch fire	///	reacts very quickly
///	reacts quickly	/	reacts	•••	slow or partial reaction
X	no reaction				

Exploring Science 9g Summary Sheets

Didier Musso



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