



QUESTION 10: Equilibrium and Le Chatelier's Principle (10 marks)

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QUESTION 10: Equilibrium and Le Chatelier's Principle (10 marks)

The reaction between hydrogen gas and carbon dioxide gas forms water and carbon monoxide gas, as shown in the following chemical equation. The system is at equilibrium at 1000 K.

$$\text{H}_2(\text{g}) + \text{CO}_2(\text{g}) \rightleftharpoons \text{H}_2\text{O}(\text{g}) + \text{CO}(\text{g})$$


- A. The concentration of the hydrogen gas, $\text{H}_2(\text{g})$, increases.
- B. The reaction between hydrogen gas and carbon dioxide gas is stopped.
- C. The rate of formation of carbon monoxide gas increases, the rate of formation of hydrogen gas decreases.
- D. The reaction stops because the system has reached equilibrium.

QUESTION 10: Equilibrium and Le Chatelier's Principle (10 marks)

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Explorelearning Equilibrium And Concentration Gizmo

S Ben Porath



Explorelearning Equilibrium And Concentration Gizmo:

This book delves into Explorelearning Equilibrium And Concentration Gizmo. Explorelearning Equilibrium And Concentration Gizmo is a crucial topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Explorelearning Equilibrium And Concentration Gizmo, encompassing both the fundamentals and more intricate discussions.

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- Chapter 1: Introduction to Explorelearning Equilibrium And Concentration Gizmo
- Chapter 2: Essential Elements of Explorelearning Equilibrium And Concentration Gizmo
- Chapter 3: Explorelearning Equilibrium And Concentration Gizmo in Everyday Life
- Chapter 4: Explorelearning Equilibrium And Concentration Gizmo in Specific Contexts
- Chapter 5: Conclusion

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