

Name: \_\_\_\_\_

## Balancing Chemical Equations



Directions: Balance each of the following chemical equations.

- 1)  $\underline{\quad} \text{H}_3\text{PO}_4 + \underline{\quad} \text{KOH} \rightarrow \underline{\quad} \text{K}_3\text{PO}_4 + \underline{\quad} \text{H}_2\text{O}$
- 2)  $\underline{\quad} \text{K} + \underline{\quad} \text{B}_2\text{O}_3 \rightarrow \underline{\quad} \text{K}_2\text{O} + \underline{\quad} \text{B}$
- 3)  $\underline{\quad} \text{HCl} + \underline{\quad} \text{NaOH} \rightarrow \underline{\quad} \text{NaCl} + \underline{\quad} \text{H}_2\text{O}$
- 4)  $\underline{\quad} \text{Na} + \underline{\quad} \text{NaNO}_3 \rightarrow \underline{\quad} \text{Na}_2\text{O} + \underline{\quad} \text{N}_2$
- 5)  $\underline{\quad} \text{C} + \underline{\quad} \text{S}_8 \rightarrow \underline{\quad} \text{CS}_2$
- 6)  $\underline{\quad} \text{Na} + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{Na}_2\text{O}$
- 7)  $\underline{\quad} \text{N}_2 + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{N}_2\text{O}_5$
- 8)  $\underline{\quad} \text{H}_3\text{PO}_4 + \underline{\quad} \text{Mg}(\text{OH})_2 \rightarrow \underline{\quad} \text{Mg}_3(\text{PO}_4)_2 + \underline{\quad} \text{H}_2\text{O}$
- 9)  $\underline{\quad} \text{NaOH} + \underline{\quad} \text{H}_2\text{CO}_3 \rightarrow \underline{\quad} \text{Na}_2\text{CO}_3 + \underline{\quad} \text{H}_2\text{O}$
- 10)  $\underline{\quad} \text{KOH} + \underline{\quad} \text{HBr} \rightarrow \underline{\quad} \text{KBr} + \underline{\quad} \text{H}_2\text{O}$
- 11)  $\underline{\quad} \text{Na} + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{Na}_2\text{O}$
- 12)  $\underline{\quad} \text{Al}(\text{OH})_3 + \underline{\quad} \text{H}_2\text{CO}_3 \rightarrow \underline{\quad} \text{Al}_2(\text{CO}_3)_3 + \underline{\quad} \text{H}_2\text{O}$
- 13)  $\underline{\quad} \text{Al} + \underline{\quad} \text{S}_8 \rightarrow \underline{\quad} \text{Al}_2\text{S}_3$
- 14)  $\underline{\quad} \text{Cs} + \underline{\quad} \text{N}_2 \rightarrow \underline{\quad} \text{Cs}_3\text{N}$
- 15)  $\underline{\quad} \text{Mg} + \underline{\quad} \text{Cl}_2 \rightarrow \underline{\quad} \text{MgCl}_2$
- 16)  $\underline{\quad} \text{Rb} + \underline{\quad} \text{RbNO}_3 \rightarrow \underline{\quad} \text{Rb}_2\text{O} + \underline{\quad} \text{N}_2$
- 17)  $\underline{\quad} \text{C}_6\text{H}_6 + \underline{\quad} \text{O}_2 \rightarrow \underline{\quad} \text{CO}_2 + \underline{\quad} \text{H}_2\text{O}$
- 18)  $\underline{\quad} \text{N}_2 + \underline{\quad} \text{H}_2 \rightarrow \underline{\quad} \text{NH}_3$

# Explorelearning Chemical Equations Answers

**Rumi Michael Leigh**



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