

## Student Exploration: Covalent Bonds

<p><b>Activity B:</b> <b>Building molecules</b></p>	<p><u>Get the Gizmo ready:</u></p> <ul style="list-style-type: none"> <li>• Click <b>Reset</b>.</li> <li>• Turn off <b>Show Lewis diagram</b>.</li> <li>• Select <b>Oxygen</b>.</li> </ul>	
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**Question:** How do atoms share more than one pair of electrons?

- Observe:** Like fluorine and most other elements, oxygen atoms are most stable with a full complement of eight valence electrons.
  - How many valence electrons does each oxygen atom have now? 6
  - How many more electrons does each oxygen atom need to be stable? 2
- Form a bond:** Drag electrons back and forth until the molecule of oxygen ( $O_2$ ) is stable. Click **Check** to confirm your molecule is stable. Take a snapshot and paste the image into your document (don't forget to label it " $O_2$ ").

How many *pairs* of shared electrons are there in a stable molecule of oxygen? 1

- Draw a diagram:** Draw a Lewis diagram of the oxygen molecule in the space below at left. To check your work, turn on **Show Lewis diagram**. Draw the correct diagram on the right.

Practice diagram:  $\cdot\ddot{O} = \ddot{O}\cdot$

Actual:  $\cdot\ddot{O} = \ddot{O}\cdot$

- Practice:** Create covalent bonds and stable molecules for the remaining substances. Take a snapshot of each completed molecule and add it to your document. Draw Lewis diagrams for each one. (As above, draw the diagram on your own before checking your work.)

Nitrogen



Ammonia



Chlorine



Methane



Water



Carbon dioxide



Formaldehyde



(Activity B continued on next page)

# Explore Learning Covalent Bonds Answer Key

**ML Yell**



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