

Iron Filings

A Mole Lab That Actually Works!



Introduction

Students will grasp the meaning of a mole and stoichiometry after performing this simple experiment.

Concepts

- Mole relations
- Oxidation–reduction

Background

Iron powder will react with copper(II) sulfate in a one-to-one ratio (1 mole to 1 mole). The students will also be able to more easily visualize the size of a mole of iron and copper. The results obtained in this lab usually have less than a one-percent error.



In the reaction, iron will be the limiting reagent; it will be completely used up in the reaction. Copper(II) sulfate will be in excess; not all of it will be used up in the reaction. The number of moles of iron that react will equal the number of moles of copper produced. The reaction indicates the 1:1 ratio between the iron and copper.

The occurrence of the reaction will be obvious since the blue copper(II) sulfate solution turns green due to the iron going into the solution. Copper metal is noticeably forming on the bottom. If the copper(II) sulfate solution is warm, the reaction occurs almost instantaneously.

This lab can also be used to illustrate a single replacement reaction and oxidation–reduction. As a single replacement reaction, iron replaces the copper in copper(II) sulfate to produce iron(II) sulfate and copper metal. As an oxidation–reduction reaction, the copper in copper(II) sulfate is reduced and the iron powder is oxidized. (Copper goes from a +2 oxidation state to a zero oxidation state while iron goes from a zero oxidation state to a +2 oxidation state.)

Materials

- Copper(II) sulfate pentahydrate, $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$, 12.5 g
- Iron powder, Fe, 2.24 g
- Water, distilled or deionized, approximately 100 mL
- Balance, accurate to 0.01 g
- Beaker, 100-mL
- Beaker, 250-mL

- Graduated cylinder, 50-mL, or 100-mL
- Glass stirring rod
- Hot plate
- Oven or heat lamp (optional)
- Weighting paper or weighing dish
(a 3" × 3" piece of paper will do)

Safety Precautions

Copper(II) sulfate (pentahydrate) is a skin and respiratory irritant. It is toxic by ingestion and inhalation (LD₅₀ 300 mg/kg). Students should be warned not to ingest any of the materials, to use them only in the manner for which they are intended, and to wash their hands thoroughly after use. Wear chemical splash goggles, chemical-resistant gloves, and a chemical-resistant apron. Please review current Material Safety Data Sheets for additional safety, handling, and disposal information.

Procedure

1. Mass a clean, dry, 100-mL beaker. Record the mass in a data table.
2. Weigh out 12.5 g of copper(II) sulfate in the previously massed 100-mL beaker.
3. Using the graduated cylinder, measure out 50.0 mL of distilled water, and then add the water to the 100-mL beaker.

Mole Lab Iron Filings Copper Chloride Answers

S Marginson



Mole Lab Iron Filings Copper Chloride Answers:

This book delves into Mole Lab Iron Filings Copper Chloride Answers. Mole Lab Iron Filings Copper Chloride Answers is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Mole Lab Iron Filings Copper Chloride Answers, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
 - Chapter 1: Introduction to Mole Lab Iron Filings Copper Chloride Answers
 - Chapter 2: Essential Elements of Mole Lab Iron Filings Copper Chloride Answers
 - Chapter 3: Mole Lab Iron Filings Copper Chloride Answers in Everyday Life
 - Chapter 4: Mole Lab Iron Filings Copper Chloride Answers in Specific Contexts
 - Chapter 5: Conclusion
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