

## Chapter 5 Chapter Review Question Answers

- Describe the scientists' contributions:
  - Cannizzaro—developed a standard method for measuring atomic masses, which allowed chemists to search for periodic trends among elements.
  - Mendeleev—organized elements according to increasing atomic mass and noticed that similar properties appeared periodically.
  - Moseley—discovered that atomic number, not atomic mass, should be the basis of organizing elements on the periodic table.
- The physical and chemical properties of the elements are periodic functions of their atomic numbers.
- Groups of elements exhibit similar chemical and physical properties and behaviors.
- Electron configurations:
  - Generally, the configurations of the outermost electrons shells of elements within the same group are the same. There are a number of exceptions to this rule, however, among the transition elements.
  - Their outer shells are completely filled.
- The length of a period is determined by the total number of electrons that can fill the outer sublevels of the elements of that period.
- An element's period corresponds to its highest occupied main energy level. (s & p)
- Blocks & Groups
  - The type of sublevel being filled in successive elements of that block.
  - s-block: Groups 1 & 2; p-block: Groups 13–18; d-block: Groups 3–12; f-block: No group #s, as they are between groups 3 & 4, but period 6 elements are in the Lanthanide series, and period 7 elements are in the Actinide series.
- Alkali Metals
  - The Group 1 elements, except Hydrogen.
  - Alkali metals are extremely reactive, they react violently with water, they have low densities, they are soft enough to be cut with a knife, and freshly cut alkali metals are very shiny silver colored metals.
- Alkaline Metals
  - The Group 2 elements.
  - They are slightly less reactive than Group 1 elements; they are harder, duller, denser than Group 1 elements. They also have slightly higher melting points than the Group 1 elements.
- Group Configurations:
  - Transition Metals (d-block)
    - Group 3:  $ns^2 (n-1)d^1$
    - Group 4:  $ns^2 (n-1)d^2$
    - Group 5:  $ns^2 (n-1)d^3$
    - Group 6:  $ns^2 (n-1)d^4$  {actually is  $ns^1 (n-1)d^5$  for Cr & Mo}
    - Group 7:  $ns^2 (n-1)d^5$
    - Group 8:  $ns^2 (n-1)d^6$
    - Group 9:  $ns^2 (n-1)d^7$
    - Group 10:  $ns^2 (n-1)d^8$
    - Group 11:  $ns^2 (n-1)d^9$  {actually is  $ns^1 (n-1)d^{10}$  for all group 11 elements}
    - Group 12:  $ns^2 (n-1)d^{10}$
  - The group number is the sum of the s & d electrons.
- The transition elements.
- p-block
  - The p-block consists of nonmetals at the right, metalloids in the middle along the staircase, and metals at the left.
  - The p-block metals are generally harder and denser than the s-block metals, but softer and less dense than the d-block metals.
- Halogens
  - The Group 17 elements.

# Modern Chemistry Chapter 5 Review Answer Key

**ML Morrison**



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