

CHAPTER TEST

Time : 1 1/2 hours

Maximum Marks : 50

PART A : Give Short Answers :

- Q.1. Define specific heat of a substance.
Q.2. What determines the state of a substance, whether it will be solid, liquid or gas ?
Q.3. Why does a gas exert pressure ?
Q.4. What is evaporation ?
Q.5. What do you mean by fluidity ?

PART B : Give Reasons for the Following :

- Q.6. Why do we see water droplets on the outer surface of a glass containing ice-cold water ?
Q.7. Why should we wear cotton clothes in summer ?
Q.8. Why does a desert cooler cool better on a hot dry day ?
Q.9. Why are we able to sip hot tea or milk faster from a saucer rather than a cup ?
Q.10. A gas fills completely the vessel in which it is kept ?

PART C : Differentiate Between the Following :

- Q.11. How evaporation differs from boiling ?
Q.12. Differentiate between solids, liquids and gases with respect to : (a) fluidity (b) kinetic energy (c) density (d) shape (e) rigidity (f) compressibility.
Q.13. Differentiate between gas and vapour.

PART D : Give Answers in Detail :

- Q.14. (a) What produces more severe burns, boiling water or steam ?
(b) For any substance, why does the temperature remain constant during the change of state?
Q.15. (a) Give two characteristic properties of a liquid.

(b) Why does our palm feel cold when we put alcohol or perfume on it ?

(c) Give two factors which determine the rate of diffusion of a liquid in another liquid.

Q.16. What are the factors which affect the rate of evaporation ? Explain.

Q.17. What are the characteristics of the particles of matter ?

Q.18. Give Reasons Only :

(a) Water at room temperature is a liquid.

(b) Naphthalene balls disappear with time without leaving any solid.

(c) We can get small of perfume sitting several metres away.

PART E : Diagrammatic Questions :

Q.19. Show diagrammatically, the interconversion of the three states of matter.

Q.20. How will you show that, by applying pressure, particles of matter can be brought close together ?

Q.21. Draw a labelled diagram to show the sublimation of Ammonium chloride.

PART F : Observation Based Questions :

Q.22. Water is kept in an earthen pot (matka) during summer. Why is it so ?

Q.23. 1 mL of diesel is added to a beaker containing 500 mL of water and stirred. Note four observations that you can make.

Q.24. Ice melts by absorbing heat. When it melts, its temperature remains the same. Where does the heat energy go ?

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Rachel Sandford



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