

**SOLUTION**

**IDENTIFY and SET UP:** This problem involves the properties of the Carnot cycle and those of an ideal gas. We are given the number of moles  $n$  and the pressure and temperature at point  $a$  (which is at the higher of the two reservoir temperatures); we can find the volume at  $a$  using the ideal-gas equation  $pV = nRT$ . We then find the pressure and volume at points  $b$ ,  $c$ , and  $d$  from the known doubling of volume in step  $a \rightarrow b$ , from equations given in this section, and from  $pV = nRT$ . In each step we use Eqs. (20.10) and (20.11) to find the heat flow and work done and Eq. (19.13) to find the internal energy change.

**EXECUTE:** The volume at point  $a$  is  

$$V_a = \frac{nRT_a}{p_a} = \frac{(0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(500 \text{ K})}{(1.00 \times 10^5 \text{ Pa})} = 8.31 \times 10^{-3} \text{ m}^3$$

The volume at point  $b$  is twice that at point  $a$ , so  

$$V_b = 2V_a = 1.66 \times 10^{-2} \text{ m}^3$$

The pressure at point  $b$  is  

$$p_b = \frac{nRT_b}{V_b} = \frac{(0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(500 \text{ K})}{1.66 \times 10^{-2} \text{ m}^3} = 5.00 \times 10^4 \text{ Pa}$$

The volume at point  $c$  is  

$$V_c = \frac{nRT_c}{p_c} = \frac{(0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(300 \text{ K})}{(1.00 \times 10^5 \text{ Pa})} = 4.99 \times 10^{-3} \text{ m}^3$$

The volume at point  $d$  is twice that at point  $c$ , so  

$$V_d = 2V_c = 9.98 \times 10^{-3} \text{ m}^3$$

The pressure at point  $d$  is  

$$p_d = \frac{nRT_d}{V_d} = \frac{(0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(300 \text{ K})}{9.98 \times 10^{-3} \text{ m}^3} = 3.15 \times 10^4 \text{ Pa}$$

The pressure at point  $a$  is  

$$p_a = \frac{nRT_a}{V_a} = \frac{(0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(500 \text{ K})}{8.31 \times 10^{-3} \text{ m}^3} = 1.00 \times 10^5 \text{ Pa}$$

For the adiabatic expansion  $b \rightarrow c$ ,  $Q_{bc} = 0$ . From the first law of thermodynamics,  $\Delta U_{bc} = Q_{bc} - W_{bc} = -W_{bc}$ ; the work  $W_{bc}$  done by the gas in this adiabatic expansion equals the negative of the change in internal energy of the gas. From Eq. (19.13) we have  $\Delta U = nC_V\Delta T$ , where  $\Delta T = T_c - T_b$ . Using  $C_V = 20.8 \text{ J/mol}\cdot\text{K}$  for an ideal diatomic gas, we find

$$\begin{aligned} W_{bc} &= -\Delta U_{bc} = -nC_V(T_c - T_b) = nC_V(T_b - T_c) \\ &= (0.200 \text{ mol})(20.8 \text{ J/mol}\cdot\text{K})(500 \text{ K} - 300 \text{ K}) = 832 \text{ J} \end{aligned}$$

For the isothermal compression  $c \rightarrow d$ ,  $\Delta U_{cd} = 0$ ; Eq. (20.11)

$$\begin{aligned} W_{cd} &= -\int_{V_c}^{V_d} p \, dV = -\int_{V_c}^{V_d} \frac{nRT}{V} \, dV \\ &= -nRT \ln \frac{V_d}{V_c} = -(0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(300 \text{ K}) \ln \frac{9.98 \times 10^{-3} \text{ m}^3}{4.99 \times 10^{-3} \text{ m}^3} \\ &= -346 \text{ J} \end{aligned}$$

The work done by the gas in the isothermal expansion  $d \rightarrow a$  is  

$$W_{da} = -W_{cd} = 346 \text{ J}$$

The work done by the gas in the isothermal compression  $a \rightarrow b$  is  

$$W_{ab} = -\int_{V_a}^{V_b} p \, dV = -\int_{V_a}^{V_b} \frac{nRT}{V} \, dV = -nRT \ln \frac{V_b}{V_a}$$

$$\begin{aligned} W_{ab} &= -(0.200 \text{ mol})(8.31 \text{ J/mol}\cdot\text{K})(500 \text{ K}) \ln \frac{1.66 \times 10^{-2} \text{ m}^3}{8.31 \times 10^{-3} \text{ m}^3} \\ &= -576 \text{ J} \end{aligned}$$

The work done by the gas in the adiabatic compression  $d \rightarrow c$  is  

$$W_{dc} = -W_{cd} = 346 \text{ J}$$

The net work done by the gas in the cycle is  

$$W_{\text{net}} = W_{bc} + W_{cd} + W_{da} + W_{ab} = 832 \text{ J} - 346 \text{ J} + 346 \text{ J} - 576 \text{ J} = 256 \text{ J}$$

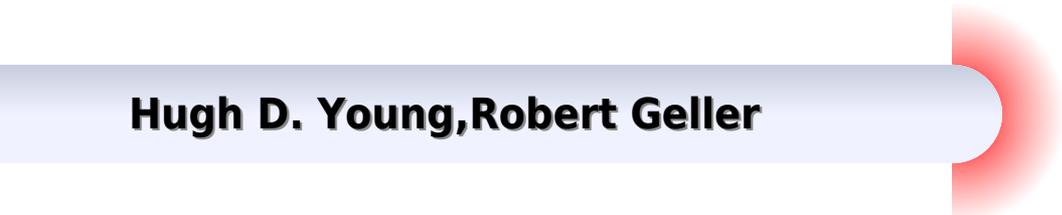
**Now Work PROBLEM 19**

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$$W_{\text{net}} = W_{bc} + W_{cd} + W_{da} + W_{ab} = 256 \text{ J}$$

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