



# SDK Installation Guide

## Java

# Java Ee Sdk Installation Guide

**Daniela Niemeyer**



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web solution 1 the two half reactions balanced as if in acidic solution  $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} + 2\text{H}^{+} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{CN}^{-}$  make electrons equal convert to basic solution  $4\text{CN}^{-} + 2\text{Au} + 2\text{CN}^{-} + 2\text{e}^{-}$  multiplied by a factor of 2  $2\text{e}^{-} + 2\text{H}_2\text{O} + \text{O}_2 \rightarrow \text{H}_2\text{O}_2 + 2\text{OH}^{-}$  3 the final answer

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being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance ce o by adding ce h<sub>2</sub>o

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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reactions  $6e$   $14h$   $Cr_2O_7^{2-}$   $2Cr^{3+}$   $7H_2O$   $Fe^{2+}$   $Fe^{3+}$   $e^-$   $2$  equalize the electrons  $6e$   $14h$   $Cr_2O_7^{2-}$   $2Cr^{3+}$   $7H_2O$   $6Fe^{2+}$   $6Fe^{3+}$   $6e^-$  multiplied by a factor of 6

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively  $C_6H_8O_6$   $2H_2NO_2$   $C_6H_6O_6$   $2H_2O_2$   $NO$  choose 1 answer  $C_6H_8O_6$  and  $NO_2$  a  $C_6H_8O_6$  and  $NO_2$  b  $NO_2$  and  $C_6H_8O_6$  b  $NO_2$  and  $C_6H_8O_6$  h and  $NO_2$  c h and

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