

How Substances Dissolve

KEY IDEAS

As you read this section, keep these questions in mind:

- Why is water called the universal solvent?
- How do substances dissolve?

Why Do Substances Dissolve in Certain Solvents?

Water is often called a universal solvent because many substances dissolve in it. However, no one substance can dissolve every solute. A general rule in chemistry is that “like dissolves like.” This rule means that a solvent will dissolve substances that have similar molecular structures. 

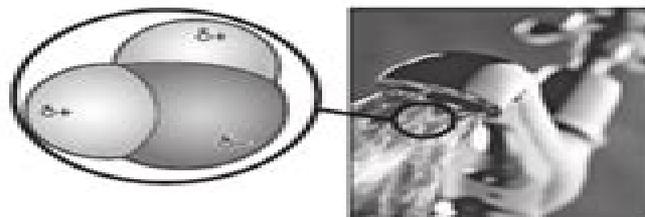
A solvent must be able to attract solute particles and pull them away from one another. Thus, solvent particles must attract solute particles more strongly than the solute particles attract one another.

POLAR MOLECULES

Recall that atoms bonded together do not all share electrons equally. For example, a water molecule is made up of two hydrogen atoms bonded to one oxygen atom. The oxygen atom attracts electrons more strongly than the hydrogen atoms do. Thus, the electrons spend more time near the oxygen atom. This gives each hydrogen atom a partial positive charge (δ^+) and the oxygen atom a partial negative charge (δ^-). 

Molecules that have partially charged positive and negative areas are **polar**. Because water molecules are polar, water can dissolve many other polar substances. Water can also dissolve ionic compounds such as table salt.

Water is a polar molecule because the oxygen atom strongly attracts electrons. Because the electrons spend more time near the oxygen atom, the oxygen atom has a partial negative charge.



READING TOOLBOX

Summarize Read each page of the section silently to yourself. With a partner, take turns summarizing each page.

READING CHECK

1. Identify Why is water called a universal solvent?

because
many substances dissolve
in it

READING CHECK

2. Explain Why do water molecules have partial charges?

The oxygen atom attracts
electrons more strongly
than the hydrogen atoms
do. Thus, the electrons
spend more time near the
oxygen atom

How Substances Dissolve Answer Key

Arthur Edward Waite



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Matter ,2004 Everything in the universe regardless of its size shape color or physical state is made up of matter

The Hermetic Museum, Restored and Enlarged ,1893 The Hermetic Museum, Restored and Enlarged Arthur Edward Waite, 1893 *Therapeutic Gazette* William Brodie, Francis Edward Stewart, John J. Mulheron, Albert Brown Lyons, 1884 *Introductory Chemistry* Charles H. Corwin, 2011 To help students learn chemical skills and concepts more effectively Introductory Chemistry Concepts and Critical Thinking Sixth Edition highlights the connection between key concepts and key problem solving skills through critical thinking Math and problem solving are covered early in the text Corwin builds your problem solving ability through innovative learning aids and technology formulated to meet your needs This revision retains all the strengths of the previous editions while adding emphasis on conceptual understanding and critical thinking Exploring Planet Earth ,1994

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Concept Learning and Problem Solving John Francis Cullen, 1983 **Process Oriented Guided Inquiry Learning (POGIL)** Richard Samuel Moog, 2008 POGIL is a student centered group learning pedagogy based on current learning theory This volume describes POGIL s theoretical basis its implementations in diverse environments and evaluation of student outcomes *General Science* ,1989 Instructors Resource Manual John McMurry, Robert A. Pribush, Robert Fay, 2000-12-27

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Table of Contents How Substances Dissolve Answer Key

1. Understanding the eBook How Substances Dissolve Answer Key
 - The Rise of Digital Reading How Substances Dissolve Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying How Substances Dissolve Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an How Substances Dissolve Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from How Substances Dissolve Answer Key
 - Personalized Recommendations
 - How Substances Dissolve Answer Key User Reviews and Ratings
 - How Substances Dissolve Answer Key and Bestseller Lists
5. Accessing How Substances Dissolve Answer Key Free and Paid eBooks
 - How Substances Dissolve Answer Key Public Domain eBooks
 - How Substances Dissolve Answer Key eBook Subscription Services
 - How Substances Dissolve Answer Key Budget-Friendly Options
6. Navigating How Substances Dissolve Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - How Substances Dissolve Answer Key Compatibility with Devices
 - How Substances Dissolve Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of How Substances Dissolve Answer Key
 - Highlighting and Note-Taking How Substances Dissolve Answer Key
 - Interactive Elements How Substances Dissolve Answer Key
 8. Staying Engaged with How Substances Dissolve Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers How Substances Dissolve Answer Key
 9. Balancing eBooks and Physical Books How Substances Dissolve Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection How Substances Dissolve Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine How Substances Dissolve Answer Key
 - Setting Reading Goals How Substances Dissolve Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of How Substances Dissolve Answer Key
 - Fact-Checking eBook Content of How Substances Dissolve Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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