

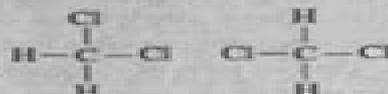
Experiment 13

The Geometrical Structure of Molecules — An Experiment Using Molecular Models

Many years ago it was observed that in many of its compounds the carbon atom formed four chemical linkages to other atoms. As early as 1870, structural formulas of carbon compounds were drawn as shown:



Although such drawings as these imply that the atom-atom linkages, indicated by valence strokes, lie in a plane, chemical evidence, particularly the existence of only one substance with the chemical formula CH_2Cl_2 , requires that the linkages be directed toward the corners of a tetrahedron, at the center of which is the carbon atom. Were the molecules truly flat, one would expect two distinct (geometric) isomers to exist, with the structural formulae:



The physical significance of the chemical linkages between atoms, expressed by the lines or valence strokes in molecular structure diagrams, became evident soon after the discovery of the electron. In a classic paper, G. N. Lewis suggested, on the basis of chemical evidence, that the single bonds in structural formulas involve two electrons and that an atom tends to hold eight electrons in its outermost or valence shell.

Lewis's proposal that atoms generally have eight electrons in their outer shells proved to be extremely useful and has come to be known as the octet rule. It can be applied to many atoms, but is particularly important in the treatment of covalent compounds of atoms in the second row of the Periodic Table. For atoms such as carbon, oxygen, nitrogen, and fluorine, the eight valence electrons occur in pairs that occupy tetrahedral positions around the central atom core. Some of the electron pairs do not participate directly in chemical bonding and are called unshared or nonbonding pairs; however, the structures of compounds containing such unshared pairs reflect the tetrahedral arrangement of the four pairs of valence shell electrons. In the H_2O molecule, which obeys the octet rule, the four pairs of electrons around the central oxygen atom occupy essentially tetrahedral positions; there are two unshared nonbonding pairs and two bonding pairs that are shared by the O atom and the two H atoms. The $\text{H}-\text{O}-\text{H}$ bond angle is nearly but not exactly tetrahedral since the properties of shared and unshared pairs of electrons are not exactly alike.



Most molecules obey the octet rule. Essentially, all organic molecules obey the rule, and so do most inorganic molecules and ions. For species that obey the octet rule it is possible to predict electron-dot, or Lewis,

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