

Experiment 25

B. Determination of the equilibrium constant for the iodine + iodide = triiodide reaction as a function of concentration.

1. Concentration of KI in original solution

$$2.00 \times 10^{-3} \text{ M}$$

2. Concentration of I_2 in original solution

$$1.00 \times 10^{-4} \text{ M}$$

3. Equilibrium expression for reaction

$$\frac{[I_3^-]}{[I_2][I^-]}$$

4. Results and calculations (Calculate the concentrations of iodine ($[I_2]_0$) and iodide ($[I^-]_0$) that would result from the dilutions if a reaction did not occur. Using the absorption measurement for triiodide (A) and its extinction coefficient, calculate the triiodide concentration ($[I_3^-]$) that results from the achievement of equilibrium, then calculate the final concentrations of iodine ($[I_2]$) and iodide ($[I^-]$) by subtraction. Finally, calculate the value of the equilibrium constant (K_{eq}).

$$A = \epsilon bc$$

#	mL I_2	mL KI	total vol. (mL)	$[I_2]_0$ (mol/L)	$[I^-]_0$ (mol/L)	A	Temp. ($^{\circ}C$)	$[I_3^-]$ (mol/L)	$[I_2]$ (mol/L)	$[I^-]$ (mol/L)	K_{eq}	H_2O mL
1	2.00	1.00	4.00	5.0×10^{-5}	5.0×10^{-3}	0.937	23 $^{\circ}$	3.09×10^{-5}	1.91×10^{-5}	4.69×10^{-3}	7.69×10^2	1.00
2	2.00	2.00	4.00	5.0×10^{-5}	1.0×10^{-2}	0.560	23	1.85×10^{-5}	3.2×10^{-5}	8.15×10^{-2}	4.71×10^2	N/A
3	3.00	1.00	4.00	7.5×10^{-5}	5.0×10^{-3}	0.456	23	1.90×10^{-5}				N/A
4	1.00	1.00	4.00	2.5×10^{-5}	5.0×10^{-3}	0.089	23	2.90×10^{-5}				2.00
5	1.00	3.00	4.00	2.5×10^{-5}	1.5×10^{-2}	0.214	23	7.06×10^{-6}				N/A

5. Average value of K_{eq}

Equilibrium Constant Determination Of Iodide

Didier Musso



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