

## Ionic Compounds and Formula Worksheet

Write the formula for each of the following compounds.

1. Magnesium oxide	<b>MgO</b>	2. Sodium fluoride	<b>NaF</b>
3. Aluminum nitride	<b>Al<sub>2</sub>N<sub>3</sub></b>	4. Potassium sulfide	<b>K<sub>2</sub>S</b>
5. Lithium iodide	<b>LiI</b>	6. Calcium bromide	<b>CaBr<sub>2</sub></b>
7. Beryllium oxide	<b>BeO</b>	8. Nickel chloride	<b>NiCl</b>
9. Magnesium nitride	<b>Mg<sub>3</sub>N<sub>2</sub></b>	10. Aluminum sulfide	<b>Al<sub>2</sub>S<sub>3</sub></b>
11. Calcium phosphide	<b>Ca<sub>3</sub>P<sub>2</sub></b>	12. Zinc fluoride	<b>ZnF</b>
13. Sodium oxide	<b>Na<sub>2</sub>O</b>	14. Lithium chloride	<b>LiCl</b>
15. Zinc oxide	<b>ZnO</b>	16. Silver chloride	<b>AgCl</b>
17. Barium bromide	<b>BaBr<sub>2</sub></b>	18. Potassium phosphide	<b>K<sub>3</sub>P</b>
19. Silver oxide	<b>Ag<sub>2</sub>O</b>		

Write the names for the following compounds.

1. Li <sub>2</sub> O	<b>Lithium oxide</b>	2. CaF <sub>2</sub>	<b>Calcium fluoride</b>
3. AlCl <sub>3</sub>	<b>Aluminum chloride</b>	4. K <sub>3</sub> P	<b>Potassium phosphide</b>
5. MgS	<b>Magnesium sulfide</b>	6. ZnF <sub>2</sub>	<b>Zinc fluoride</b>
7. CaO	<b>Calcium oxide</b>	8. AgCl	<b>Silver chloride</b>
9. KBr	<b>Potassium bromide</b>	10. Al <sub>2</sub> S <sub>3</sub>	<b>Aluminum sulfide</b>
11. BeF	<b>Beryllium fluoride</b>	12. BaBr <sub>2</sub>	<b>Barium bromide</b>
13. Na <sub>3</sub> N	<b>Sodium nitride</b>	14. CaS	<b>Calcium sulfide</b>
15. Al <sub>2</sub> O <sub>3</sub>	<b>Aluminum oxide</b>	16. KI	<b>Potassium iodide</b>
17. Na <sub>2</sub> S	<b>Sodium sulfide</b>	18. ZnO	<b>Zinc oxide</b>
19. Mg <sub>3</sub> P <sub>2</sub>	<b>Magnesium phosphide</b>		

# Formulae Of Ionic Compounds Answer Key

**Morris Hein, Scott Pattison, Susan  
Arena, Leo R. Best**

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