



### General Chemistry formula sheet

Prefix	Symbol	Factor
Giga	G	$10^9$
Mega	M	$10^6$
Kilo	K	$10^3$
centi	C	$10^{-2}$
milli	m	$10^{-3}$
micro	$\mu$	$10^{-6}$
nano	n	$10^{-9}$
$1 \text{ mol} = 6.022 \times 10^{23} \text{ molecules}$ $1 \text{ cm}^3 = 1 \text{ mL}$		

Name	Formula	Name	Formula
Acetate	$\text{C}_2\text{H}_3\text{O}_2^-$	Phosphate	$\text{PO}_4^{3-}$
Carbonate	$\text{CO}_3^{2-}$	Ammonium	$\text{NH}_4^+$
Bicarbonate	$\text{HCO}_3^-$	Chlorite	$\text{ClO}_2^-$
Hydroxide	$\text{OH}^-$	Chlorate	$\text{ClO}_3^-$
Nitrite	$\text{NO}_2^-$	Sulfite	$\text{SO}_3^{2-}$
Nitrate	$\text{NO}_3^-$	Sulfate	$\text{SO}_4^{2-}$
Cyanide	$\text{CN}^-$		

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$

$$P_A = X_A \cdot P_{\text{total}}$$

$$c = \lambda\nu$$

$$M = \frac{\text{mol solute}}{\text{L solution}}$$

$$PV = nRT$$

$$q = mc\Delta T$$

$$\Delta H^\circ_{\text{rxn}} = \sum n_p \Delta H^\circ_f - \sum n_r \Delta H^\circ_f$$

$$FC = \# \text{ of valence } e^- - [\# \text{ of lone pairs} + \# \text{ of bonds}]$$

$$M_1V_1 = M_2V_2$$

$$X_A = \frac{\text{mol } A}{\text{total mol}}$$

$$q = n\Delta N$$

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